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Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

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Holt ChemFile: Problem-Solving Workbook 272 Titrations Titrations Chemists have many methods for determining the quantity of a substance present in a solution or other mixture. One common method is titration, in which a solution of known concentration reacts with a sample containing the substance of unknown quantity.

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## Conversions

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### **Skills Worksheet Problem Solving**

Holt ChemFile: Problem-Solving Workbook 49 Mole Concept Name Class Date • Problem Solving continued PROBLEMS INVOLVING ATOMS AND ELEMENTS Sample Problem 1 A chemist has a jar containing 388.2 g of iron filings. How many moles of iron does the jar contain? Solution ANALYZE What is given in the problem? mass of iron in grams

### **continued - PC\|MAC**

Holt ChemFile: Problem-Solving Workbook 261 pH Name Class Date Problem Solving continued Sample Problem 1 A HCl solution has a concentration of 0.0050 M. Calculate  $[\text{OH}^-]$  and  $[\text{H}_3\text{O}^+]$  for this solution. HCl is a strong acid, so assume it is 100% ion-ized. Solution ANALYZE What is given in the problem? the molarity of the HCl solution, and the fact

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Holt ChemFile: Problem-Solving Workbook 99 Stoichiometry Name Class Date Problem Solving continued Sample Problem 1 Ammonia is made industrially by reacting nitrogen and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The equation is  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ . If 4.0 mol of  $\text{H}_2$  react, how many moles of  $\text{NH}_3$  will be produced?

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Holt ChemFile: Problem-Solving Workbook 97 Stoichiometry. Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of

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moles and that they relate moles of a substance to mass by using the molar mass.

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Holt ChemFile: Problem-Solving Workbook 201 Concentration of Solutions Sample Problem 3 Determine the molal concentration of a solution containing 81.3 g of ethylene glycol, HOCH<sub>2</sub>CH<sub>2</sub>OH, dissolved in 166 g of water. Solution ANALYZE What is given in the problem? the mass of ethylene glycol dissolved, and What are you asked to find? Items Data

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