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Crystal Violet Rate Law Lab

In this experiment, crystal violet and NaOH form a complex that changes from transparent blue to colorless over time. The absorbance is measured using a spectrophotometer, and the rate law is then determined using this information.

Experimental. First, a spectrophotometer was

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turned on and set at a wavelength of 595 nm.

Determining the Rate Law for the Crystal Violet-Hydroxide ...

(crystal violet) The rate law for this reaction would then be in the form $\text{Rate} = k [\text{CV}^+]^x [\text{OH}^-]^y$

However, in order to use graphical analysis to determine reaction orders, pseudo reaction conditions are

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necessary. In this case, the reactant that will be in excess is the sodium hydroxide. Thus, the rate law can be rewritten as

Experiment 7 Rate Law Determination of the Crystal Violet

...

In this experiment you will determine the rate law for the reaction of the dye crystal violet (CV) with OH^- in aqueous solution

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according to the balanced net ionic equation given in Scheme 1. We will define the rate of reaction as the disappearance of the colored CV over time, which can be expressed in differential form as $d[CV]/dt$.

Kinetics of Crystal Violet Bleaching | Chem Lab

Studying the graphs,

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we determined that the rate was in first order with respect to Crystal Violet: Rate = $k[\text{CV}]$ 1. Moreover, using Beer's Law, we substituted our data into the standard first order equation: $\ln(\epsilon bc t) = -k(t) + \ln(\epsilon bc o)$, finding that the rate constant is approximately 0.0909. This seems a reasonable answer, but there ...

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Rate Law

Determination of a Crystal Violet Reaction

Experiment 7 Rate Law
Determination of the
Crystal Violet Reaction

OUTCOMES After
completing this
experiment, the
student should be able
to: use graphical
analysis to determine
the order of a reaction.
determine the pseudo
rate constant and half-
life for a reaction. write

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Comprehending as
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can be taken as with
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[MOBI] Crystal Violet Rate Law Lab

In this investigation, we will derive the rate law for the decolorization of crystal violet by hydroxide. In order to determine the rate law, we need to design an experiment that measures the concentration of a species at a particular time during a reaction.

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Lab Investigation 4 - How Fast Does the Crystal Violet ...

Write the correct rate law expression for the reaction, in terms of crystal violet only (omit OH⁻). Absorbance is proportional to the concentration of crystal violet ($A = \epsilon l [CV^+]$) and can be used instead of concentration when plotting data ($A \approx [CV^+]$). rate1 = $-\Delta[CV^+]/\Delta t =$

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$k_1[CV^+]^m$ where $k_1 =$
 $k[OH^-]^{1n}$; $[OH^-]$ is
0.020 M rate₁ =
-0.0821[CV⁺]¹

RATE LAW DETERMINATION OF CRYSTAL VIOLET HYDROXYLATION ...

To find the rate law of the reaction between crystal violet and sodium hydroxide, the rate constant needs to be determined from the linear graph. In a first order reaction, the

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slope of the ln...

Chemical Kinetics: Finding the Rate Law (Kathryn Smith

...

Differential rate law for hydroxylation of crystal violet. rate =

$-\Delta[\text{CV}^+] / \Delta(t)$

rate =

$k[\text{CV}^+]^m[\text{OH}^-]^n$. k

= rate constant. m =

order with respect to

crystal violet. n = order

with respect to

hydroxide ion. Pseudo

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rate constants. in both trials, hydroxide ion is huge excess and can be assumed constant.

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Flashcards | Quizlet

The reaction between crystal violet and hydroxide ion is presented below: The rate expression for this reaction is $\text{rate} = k [\text{CV}^+]^m [\text{OH}^-]^n$

(2) Where k = rate constant, m is the

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order of reaction with respect to CV^+ and n is the order of reaction with OH^- .

Rate Law

Determination of the Crystal Violet Reaction ...

Crystal violet is a common, beautiful purple dye. In strongly basic solutions, the bright color of the dye slowly fades and the solution becomes colorless. The kinetics

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Lab #11 - Kinetics of Crystal Violet Fading - LHS AP Chemistry

An Advanced Inquiry

Lab Introduction

Crystal violet is a common, beautiful purple dye. In strongly basic solutions, the bright color of the dye slowly fades and the solution becomes colorless. The kinetics of this "fading"

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reaction can be analyzed by measuring the color intensity or absorbance of the solution versus time to determine the rate law.

Catalog No. AP7644 **Publication No. 7644** **Kinetics of ...**

The kinetics of this reaction can be monitored with a spectrophotometer by observing the decrease in absorbance of crystal violet with time.

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The rate law in general form is: rate of disappearance of CV = rate of appearance of CVOH = $k [CV]^x [OH^-]^y$
(1)

Experiment 6: Chemical Kinetics

Title: Experiment 7

Author: Lance S. Lund

Created Date:

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Experiment 7 - Anoka-Ramsey Community College

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0.005M Sodium hydroxide, 6.75×10^{-6} M crystal violet for first run of the experiment. In the second run, replace 0.005M sodium hydroxide with 0.01M sodium hydroxide. Use the questions given below to guide you write a good report.

Lab report for Chemistry(Reaction between Crystal Violet ...

The determination of

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the rate law for the reaction of crystal violet with hydroxide is completed in this experiment. The order of the reaction with respect to CV is determined by comparison with the integrated rate laws for first and second order reactions. The rate constant is also determined.

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