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reaction rate equation.  
instantaneous rate.  
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rate at which reactants  
change into products  
over time. change in  
concentration/change  
in time. The reaction  
rate at a particular  
time as opposed to the  
avera....  $aA + bB = cC$   
 $+dD$ .

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Answers: (a)  $8.4 \times 10^{-7}$  M/s, (b)  $2.1 \times 10^{-7}$  M/s  
SAMPLE EXERCISE

14.3 continued The decomposition of  $\text{N}_2\text{O}_5$  proceeds according to the following equation: If the rate of decomposition of  $\text{N}_2\text{O}_5$  at a particular instant in a reaction vessel is  $4.2 \times 10^{-7}$  M/s, what is the rate of appearance of (a)  $\text{NO}$

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2, (b) 0 2?

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conjugate base of  
 $\text{HSO}_4^-$  is A)  $\text{H}_2\text{SO}_4$

B)  $\text{SO}_4^{2-}$  C)  $\text{H}_3\text{SO}_4^+$

D)  $\text{HSO}_4^+$  E)  $\text{OH}^-$  2) The

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neutrons reside in the  
nucleus of atoms, they  
are called nucleons.

16.

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Equilibrium cannot be established between the liquid and the gas phase if the top is removed from the bottle because the system is not closed; one of the components of the equilibrium, the

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Br<sub>2</sub> vapor, would escape from the bottle until all liquid disappeared. Thus, more liquid would evaporate than can condense back from the gas phase to the liquid phase.

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have a mass of  
157amu, 14 have a  
mass of 158amu, 14  
have a mass of

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159amu, and the remainder have a mass of 160amu? Solve for % of 160amu isotope:  
 $100 - 2 - 12 - 14 - 14 = 58\%$   
 $(.02 \times 156) + (.12 \times 157) + (.14 \times 158) + (.14 \times 159) + (.58 \times 160) = 159.14$   
amu 5.

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